

Argonne National Laboratory

**APPLICATION OF THE VACUUM-CUP
TECHNIQUE TOWARD SPECTROCHEMICAL
ANALYSIS OF URANIUM-FISSIUM ALLOY**

by

Yu Yokoyama and John P. Faris

LEGAL NOTICE

This report was prepared as an account of Government sponsored work. Neither the United States, nor the Commission, nor any person acting on behalf of the Commission:

- A. Makes any warranty or representation, expressed or implied, with respect to the accuracy, completeness, or usefulness of the information contained in this report, or that the use of any information, apparatus, method, or process disclosed in this report may not infringe privately owned rights; or*
- B. Assumes any liabilities with respect to the use of, or for damages resulting from the use of any information, apparatus, method, or process disclosed in this report.*

As used in the above, "person acting on behalf of the Commission" includes any employee or contractor of the Commission, or employee of such contractor, to the extent that such employee or contractor of the Commission, or employee of such contractor prepares, disseminates, or provides access to, any information pursuant to his employment or contract with the Commission, or his employment with such contractor.

ARGONNE NATIONAL LABORATORY

9700 South Cass Avenue

Argonne, Illinois 60440

APPLICATION OF THE VACUUM-CUP TECHNIQUE
TOWARD SPECTROCHEMICAL ANALYSIS
OF URANIUM-FISSIUM ALLOY

by

Yu Yokoyama* and John P. Faris

Chemistry Division

*Present Address: Atomic Fuel Corporation, Tokai-Mura,
Ibaraki-Ken, Japan

July 1964

Operated by The University of Chicago
under
Contract W-31-109-eng-38
with the
U. S. Atomic Energy Commission

TABLE OF CONTENTS

	<u>Page</u>
ABSTRACT	3
I. INTRODUCTION.	3
II. APPARATUS.	4
III. REAGENTS.	4
IV. SELECTION OF ANALYTICAL LINES.	5
V. EXPERIMENTAL.	6
VI. ANALYTICAL.	11
VII. SUMMARY	12
VIII. LITERATURE CITED	13

I. INTRODUCTION

The vacuum-cup electrode technique of spectrochemical analysis was introduced by Hsieh⁽¹⁾ in 1957. During the oxidation period the sample solution is led from a Teflon cup up through an axial capillary electrode into the analytical gap. With the reduced pressure in the gap the heat of discharge contributes to the feeding mechanism, with the action related to the number of bubbles per half cycle of solvent. Hsieh and Coulter⁽²⁾ employed a vacuum-cup electrode in the analysis of various metals. Leister⁽³⁾ studied the influence of various factors on the analysis of arsenic, bismuth, copper, iron, and tin in solution by electrodeposition, and reported that the sensitivity was a function of the electrode current and the amount of graphite consumed in the electrochemical gap. The method has shown good promise for the direct analysis of the sample.

APPLICATION OF THE VACUUM-CUP TECHNIQUE TOWARD SPECTROCHEMICAL ANALYSIS OF URANIUM-FISSIUM ALLOY

by

Yu Yokoyama and John P. Faris

ABSTRACT

The vacuum-cup technique of spectrographic analysis has been studied in some detail for possible use in the analysis of uranium or uranium-plutonium alloys containing some fission product elements. The changes in spectral line intensities of molybdenum, ruthenium, rhodium, and palladium were investigated in relation to several factors such as discharge type, discharge current, form of electrode, consumption of sample solution, and consumption of electrodes.

Spectral intensities correlated with consumption of sample solution but were influenced to a large extent by the deposition of solutes on the electrode surface and to some degree by concentration of the sample solution by heat. Satisfactory working curves could be obtained by adding nickel to serve as an internal standard. For this particular application, however, the method offered no outstanding advantages over other solution techniques of spectrochemical analysis.

I. INTRODUCTION

The vacuum-cup electrode technique of spectrochemical analysis was introduced by Zink⁽⁶⁾ in 1959. During the excitation period the sample solution is fed from a Teflon cup up through an axial canal in a graphite electrode into the analytical gap. Both the reduced pressure in the gap and the heat of discharge contribute to the feeding mechanism, with the pumping action related to the number of breaks per half cycle of current.⁽⁶⁾ Bass and Soulati⁽¹⁾ employed a vacuum-cup electrode in the analysis of plant materials. Leistner⁽⁴⁾ studied the influence of various factors on the intensities of arsenic, boron, copper, iron, and tin in solution by employing uni-arc excitation, and reported that the sensitivity was a function of the discharge current and the amount of graphite consumed in the analytical gap. The method has shown good promise for the direct analysis of oil samples.⁽⁵⁾

The use of spectrochemical methods could offer an appreciable savings of time in the analysis of complex fissium alloys because fairly lengthy chemical procedures are required for determination of the individual components.⁽²⁾ In this study the vacuum-cup technique was examined in particular as a means for determination of four of the elements in a uranium matrix without prior separation. The approximate concentrations expected were molybdenum - 3.4%, ruthenium - 2.6%, rhodium - 0.5%, and palladium - 0.3%. All studies were with commercially available preformed electrodes.

II. APPARATUS

The spectrograph used for this work was a 21-ft Jarrell-Ash Wadsworth Stigmatic Grating Spectrograph, having a linear dispersion of 5 Å/mm in the first order. A Jarrell-Ash Standard Varisource was employed for excitation of the spectra.

Preformed United Carbon vacuum-cup electrodes (UCP-6011) were used together with the Teflon cup (UCP-6010-A) and National Spectroscopic counter electrode (L-4036). Other types of graphite were also used in experiments to determine the optimum size, shape, and density of the counter electrode.

Spectra were recorded on Eastman Kodak SA-1 and 103-0 spectrographic plates and the spectral line intensities measured with a Jarrell-Ash Console Microphotometer.

III. REAGENTS

Standard solutions of molybdenum (5.06 mg/ml), ruthenium (7.1 mg/ml), rhodium (10 mg/ml), palladium (10 mg/ml), cobalt (1 mg/ml) and nickel (2 mg/ml) were prepared. Synthetic solutions of known concentration (see Tables I and II) were made by mixing these solutions and diluting with distilled water, distilled hydrochloric acid (6 M), or analytical grade concentrated sulfuric acid.

Table I

SYNTHETIC SOLUTIONS A

	Rh (ppm)	Pd (ppm)	Ru (ppm)	Mo (ppm)	HCl (N)
A1	10	20	46.5	50	3
A2	20	40	93	100	3
A3	50	100	233	250	3
A4	100	200	465	500	3

Table II

SYNTHETIC SOLUTIONS B

	<u>Rh</u> <u>(ppm)</u>	<u>Ni</u> <u>(mg/ml)</u>	<u>HCl</u> <u>(N)</u>
B1	80	1	2
B2	40	1	2
B3	20	1	2
B4	10	1	2

IV. SELECTION OF ANALYTICAL LINES

The spectra of uranium and each element to be studied were photographed and observed. Several of the strongest lines having nearly the same excitation potentials and which were free from spectral interference are listed below in Table III.

Table III

SELECTED SPECTRAL LINES

	<u>Wavelength,</u> <u>Å</u>	<u>Excitation</u> <u>Potential, eV</u>	<u>Remarks</u>
Mo I	3193.97	3.88	
Mo I	3170.35	3.91	
Mo I	3132.59	3.96	
Ru I	3728.03	3.32	
Ru I	3661.35	3.53	
Ru I	3436.74	3.75	U 3436.78 Å
Rh I	3692.36	3.35	{ Pu 3435.0 Å Mo 3434.79 Å
Rh I	3434.89	3.60	
Rh I	3396.85	3.64	
Pd I	3634.70	4.23	
Pd I	3421.24	4.58	{ Pu 3421.34 Å Pu 3404.69 Å Mo 3404.34 Å
Pd I	3404.58	4.46	
Pd I	3242.70	4.64	

V. EXPERIMENTAL

A. Excitation

Adequate intensities of the lines of the four elements to be determined were not obtained by a spark method. Excitation of these elements by ignited AC arc and by ignited unidirectional arc was investigated, varying the frequency of break per half-cycle and arc-power setting. It was found that an increase in frequency of break per half-cycle at a constant arc power setting caused an increase in the consumption of sample solution. This increase was accompanied by a rise in the discharge current (see Fig. 1) and resulted in a greater intensity of both the element spectra and the OH molecular spectra. A direct correlation between consumption of graphite electrode and discharge current was not clearly indicated (see Fig. 2).

B. Counter Electrode

In some preliminary experiments, in which National electrode L-4036 was used as the counter electrode, the consumption of graphite during a 60-sec discharge was usually from 1 to 2 mm, and occasionally more than 3 mm. The discharge current decreased with increase of gap distance; for example, the discharge current was 5.8, 4.7, and 2.8 A at gap distances of 1, 2, and 3 mm, respectively. This decrease of current with time was not desirable during an analytical determination. Although Leistner⁽⁴⁾ found some correlation between spectral line intensity and counter-electrode consumption, we found that the volume of sample solution consumed had a much greater effect and strived to minimize any change in gap distances.

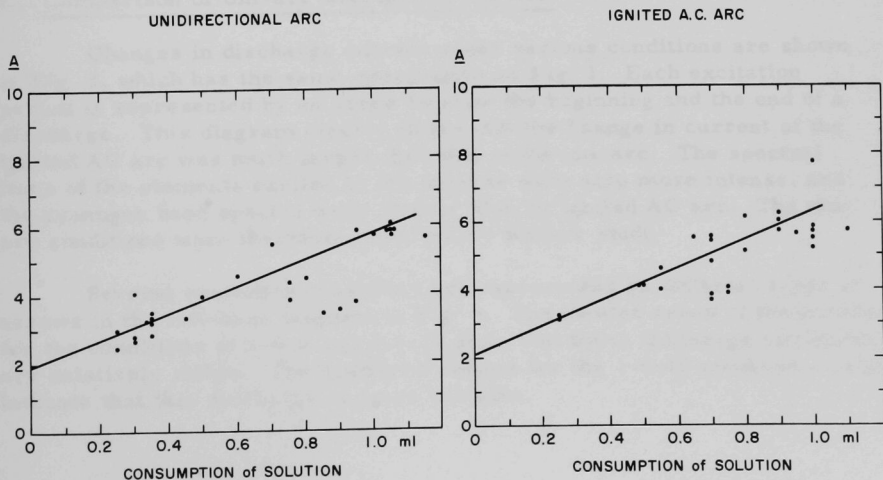


Fig. 1. Correlation between Consumption of Sample Solution and Discharge Current ($t = 60$ sec)

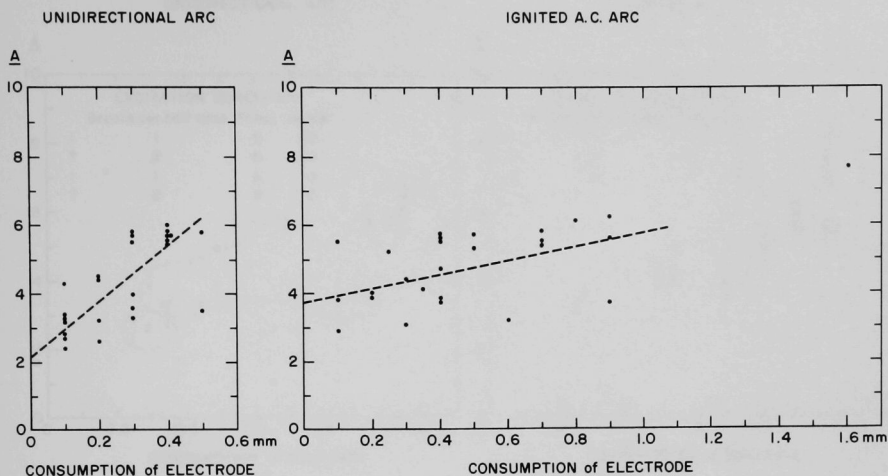


Fig. 2. Correlation between Consumption of Graphite Electrode and Discharge Current ($t = 60$ sec)

Examination of electrodes with different densities and various diameters showed that the most stable discharge was obtained with a flat-end counter electrode of about 3-mm diameter. No appreciable difference was found between regular and high-density graphite under the conditions used. A new graphite counter electrode was used for each discharge because the consumption rate tended to become erratic upon repeated use.

C. Comparison of Uni-arc with Ignited AC Arc

Changes in discharge current under various conditions are shown in Fig. 3, which has the same coordinates as Fig. 1. Each excitation period is represented by an arrow to show the beginning and the end of a discharge. This diagram clearly shows that the change in current of the ignited AC arc was much larger than that of the uni-arc. The spectral lines of the elements excited by the uni-arc were also more intense, and the cyanogen band spectra were weaker than by ignited AC arc. The uni-arc conditions were therefore selected for further study.

Several excitation conditions are represented by different types of arrows in the left-hand diagram in Fig. 3. The limited areas of the arrows for the conditions of 1-6-0 and 2-6-10 show that these discharge conditions are relatively stable. The scattered arrows for the 1-6-10 condition would indicate that this discharge is more unstable.

UNIDIRECTIONAL ARC

IGNITED A.C. ARC

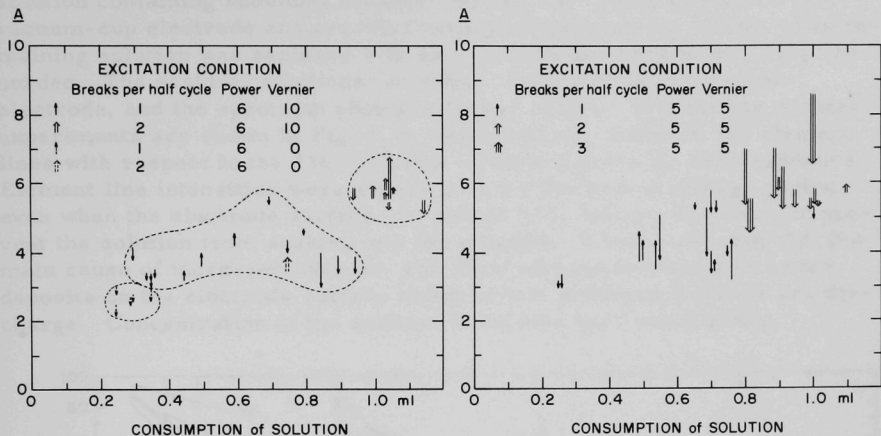


Fig. 3. Change in Current during Discharge Period (60 sec)

In the uni-arc excitation the polarity of the vacuum-cup electrode was normally positive. When the polarity was reversed, the counter electrode became very hot and rapidly vaporized. The discharge current was small, very little solution was consumed, and satisfactory results were not obtained.

D. Variation in Spectral Intensity

The standard solutions shown in Table I were excited under the above condition of 2-6-10 to examine the reproducibility of the spectral intensity of the four elements to be determined. In this experiment the same vacuum-cup electrode was used for four exposures of each standard, replenishing the solution in a cup after each discharge. The spectra were photographed for 60 sec and the transmissions measured, as shown in Fig. 4. The intensities of the OH band component at 3361.9 Å, produced by excitation of water vapor, vary only within a limited range, indicating a nearly constant feed rate of solution into the analytical gap. On the other hand, line intensities of the elements in solution were low in the beginning but increased with further use of the same electrode. During the first exposure period the spectral intensities did not show linearity, but the intensities from the the second discharge lay nearly on a straight line.

Possible reasons for the difference observed between the element spectra and the OH band spectra could be: (1) absorption of the elements by graphite, (2) deposition of solutes on the electrodes, and (3) concentration of the sample solution by heat from the discharge. An attempt was

made to determine the major cause by the following experiment: a standard solution containing rhodium, uranium, nickel, and cobalt was added to a vacuum-cup electrode and arc-discharge photographed for 60 sec. The remaining solution was replaced with new solution and another discharge recorded. The residual solutions were then combined, added to a new electrode, and the spectrum photographed as before. The results of these experiments are shown in Fig. 5, in which intensity ratios of the element lines with respect to the 3361.9-Å line of OH are given for each exposure. Element line intensities were always less for the first discharge period, even when the electrode surface was coated with Apiezon-N grease to prevent the solution from soaking into the graphite. It was concluded that the main cause of increased intensity with time was the formation of solute deposits on the electrode surface owing to heat produced from the arc discharge. Concentration of the solution could also have some effect.

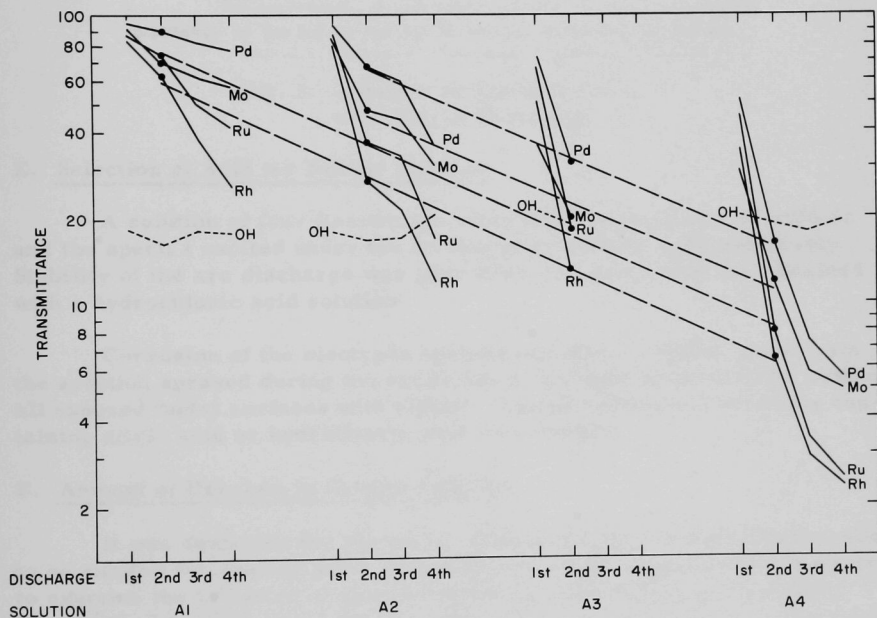


Fig. 4. Relation of Spectral Intensities with Repetition of Discharge

It has been shown in Fig. 4 that the line intensities of elements in solution tended to stabilize after an initial sparking period of 60 sec. This would indicate that neither the deposition of solutes nor their changes in concentration during arcing are an important obstacle to analytical determination by the vacuum-cup method.

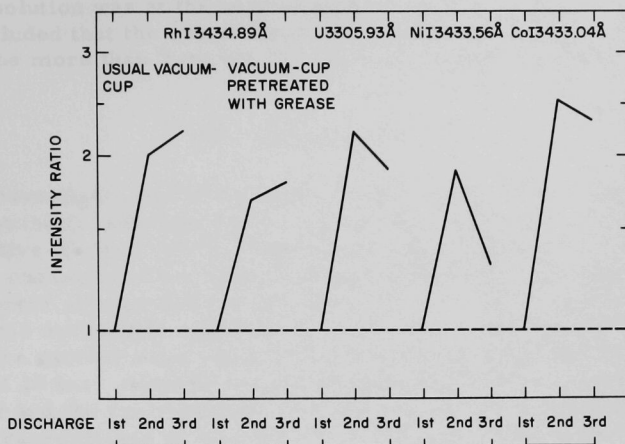


Fig. 5. Relation of Intensity Ratio with Way of Discharge

E. Selection of Acid for Sample Solution

A solution of four fissium elements was prepared in 2 N sulfuric and the spectra excited under the discharge conditions used previously. Stability of the arc discharge was poor when compared with that obtained with a hydrochloric acid solution.

Corrosion of the electrode holders and other metallic parts from the solution sprayed during the excitation period was minimized by coating all exposed metal surfaces with plastic. The use of sample solutions containing nitric acid or hydrofluoric acid was avoided.

F. Amount of Uranium in Sample Solution

It was desirable for the acidic solution of the uranium-fissium alloy to be excited without any prior separation of major constituents. In order to examine the influence of uranium concentration, solutions containing 10 mg/ml, 3 mg/ml, and 1 mg/ml of uranium were excited under the conditions described. As the uranium concentration decreased, the discharge became more stable. The discharge current was rather unsteady during excitation of the most concentrated uranium solution, consumption of sample solution was appreciably less than usual, and the background was high. During excitation of the 3-mg/ml solution the discharge was fairly stable, with the discharge current slightly lower than normal. There was no difference observed between excitation of the 1-mg/ml uranium solution and a solution containing no uranium. In both the latter cases the consumption

of sample solution was at the normal rate. From the above experiments it was concluded that the concentration of solutes in a sample solution should not be more than 3 mg/ml for the best analytical results.

VI. ANALYTICAL

To investigate analytical application of the vacuum-cup method, a set of synthetic solutions containing varying amounts of rhodium as a representative fissium element were prepared as shown in Table II. Nickel was chosen to serve as the internal standard because the excitation potentials were similar and the spectrum had a number of convenient lines located in the wavelength region of the elements to be determined. The spectra were excited under the conditions used above and photographed in intervals of 20 sec. After 80 sec the solution in a cup was replaced with a new aliquot and the run repeated. The top half of Fig. 6 shows the average change in transmittance of Rh I 3434.89 Å for a duplicate series of exposures of the 80 γ /ml standard solution. There was considerable scattering, as indicated by the height of the vertical lines on the diagram. The ratios obtained for transmittances of the rhodium line to that of Ni I 3433.56 Å are plotted for the same exposures on the bottom half of Fig. 6. Although the ratios varied with exposure time, the values for each time interval showed very little scatter and the change with time followed about the same pattern for both the first and the second use of an electrode. A similar behavior was observed also for the more dilute rhodium standard solutions.

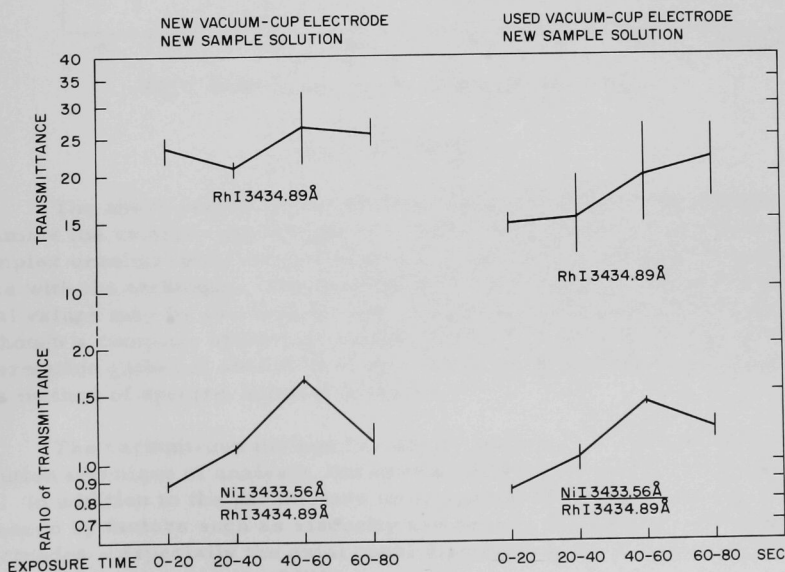


Fig. 6 Time Study of Spectral Intensity (Synthetic Solution B-1). Excitation Conditions: Igniter (Spark) $L = 310 \mu\text{H}$; $C = 0.005 \mu\text{F}$; Control Gap = 5 mm; Powerstat = 45, Current = 2 A, Break per Half-cycle = 2. Uni-arc Power = 6, Vernier = 0, Current = ~ 3.6 A.

The vacuum-cup method gave good reproducibility under constant experimental conditions, but for acceptable accuracy it was essential to photograph identical periods of sample excitation. This is further illustrated in Fig. 7 where the transmittance ratios for the standard solutions are plotted for the first period (0-20 sec) and for the third period (40-60 sec). The curves do not show linearity at the lowest concentrations because transmittance ratios were used instead of the usual intensity ratios.

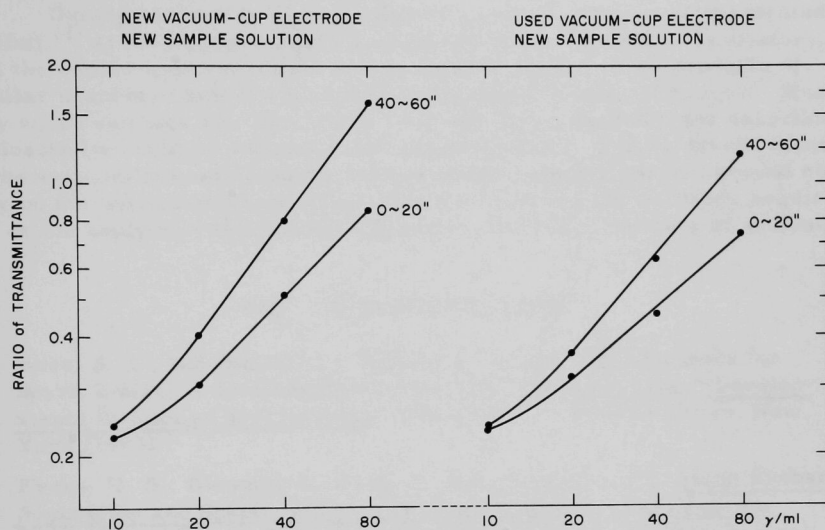


Fig. 7. Working Curve. Excitation Conditions: Same as Fig. 6

VII. SUMMARY

The above series of experiments were carried out specifically to examine the vacuum-cup method as a means for analysis of samples of complex uranium-plutonium-fission alloys and generally to gain experience with the technique. The results have shown that satisfactory analytical values may be obtained for the fission element components of an alloy. Although a complete analytical system has not been presented, most of the information gathered should be of general value toward any application of this method of spectrochemical analysis.

The vacuum-cup method has all the inherent advantages of any solution technique of analysis, but several disadvantages should be pointed out. In addition to the parameters investigated, the results may be influenced by factors such as viscosity and density of solution, variation in electrodes - especially the axial canal diameter, and the presence of any other acids, solvents, or additives.

During excitation the liquid sprayed the surrounding area. A Lucite enclosure with open ports was placed over the optical bench and connected to an exhaust system in order to contain the acid spray, but the condensing lens had to be cleaned often, and it was difficult to prevent corrosion of metallic parts. Excitation of highly radioactive samples would require an efficient, well-designed enclosure.

During the course of the present experiments it was demonstrated by Huff,⁽³⁾ at the Argonne Fuels Technology Center analytical laboratory, that the copper spark method could be readily applied to the analysis of similar uranium-fissium alloys that also contained some plutonium. Since only small samples (ca. 50 γ) were required for an analysis, the associated radioactivity could be contained with little difficulty. Future development of the vacuum-cup technique for this particular application will depend not only on the number and type of samples received and the accuracy required, but on the analytical convenience compared with other methods of analysis.

VIII. LITERATURE CITED

1. Bass, S. T., and Soulati, J., "Use of a Vacuum Cup Electrode for Spark Emission Spectrographic Analysis of Plant Tissue," Developments in Applied Spectroscopy, Vol. 2, p. 211, Plenum Press, New York (1963).
2. Evans, H. B., Bloomquist, C. A. A., and Hughes, J. P., Anion Exchange Separation and Spectrophotometric Determination of Microgram Quantities of Rhodium in Plutonium-Uranium-Fissium Alloys, Anal. Chem., 34, 1692 (1962).
3. Huff, E. A., ANL Private Communication (1963).
4. Leistner, C. J., "Spectrographic Analysis of Aqueous Solution by the Vacuum Cup Technique," 13th Annual Mid-America Spectroscopy Symposium, Chicago (1962).
5. McGowan, R. J., Spectrochemical Analysis of Oils Using "Vacuum Cup" Electrode, Appl. Spectroscopy, 15, 179 (1961).
6. Zink, T. H., A "Vacuum Cup" Electrode for the Spectrochemical Analysis of Solutions, Appl. Spectroscopy, 13, 94 (1959).

ARGONNE NATIONAL LAB WEST



3 4444 00009037 3

X