



Backend Nuclear Fuel Cycle Radiation Chemistry

October 2024

Changing the World's Energy Future

Gregory Peter Holmbeck



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**Idaho National Laboratory
Idaho Falls, Idaho 83415**

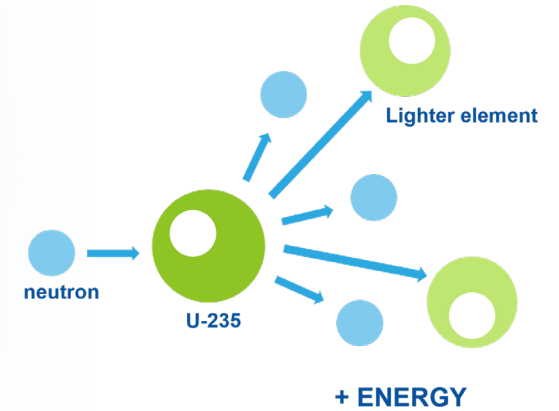
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Gregory P. Holmbeck

Center for Radiation Chemistry Research

Radiochemical Separations and Radiation Science Department



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BNL LEAF Workshop

October 21, 2024 | Brookhaven National Laboratory, NY, USA

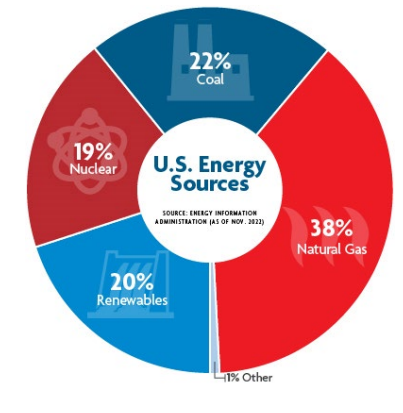
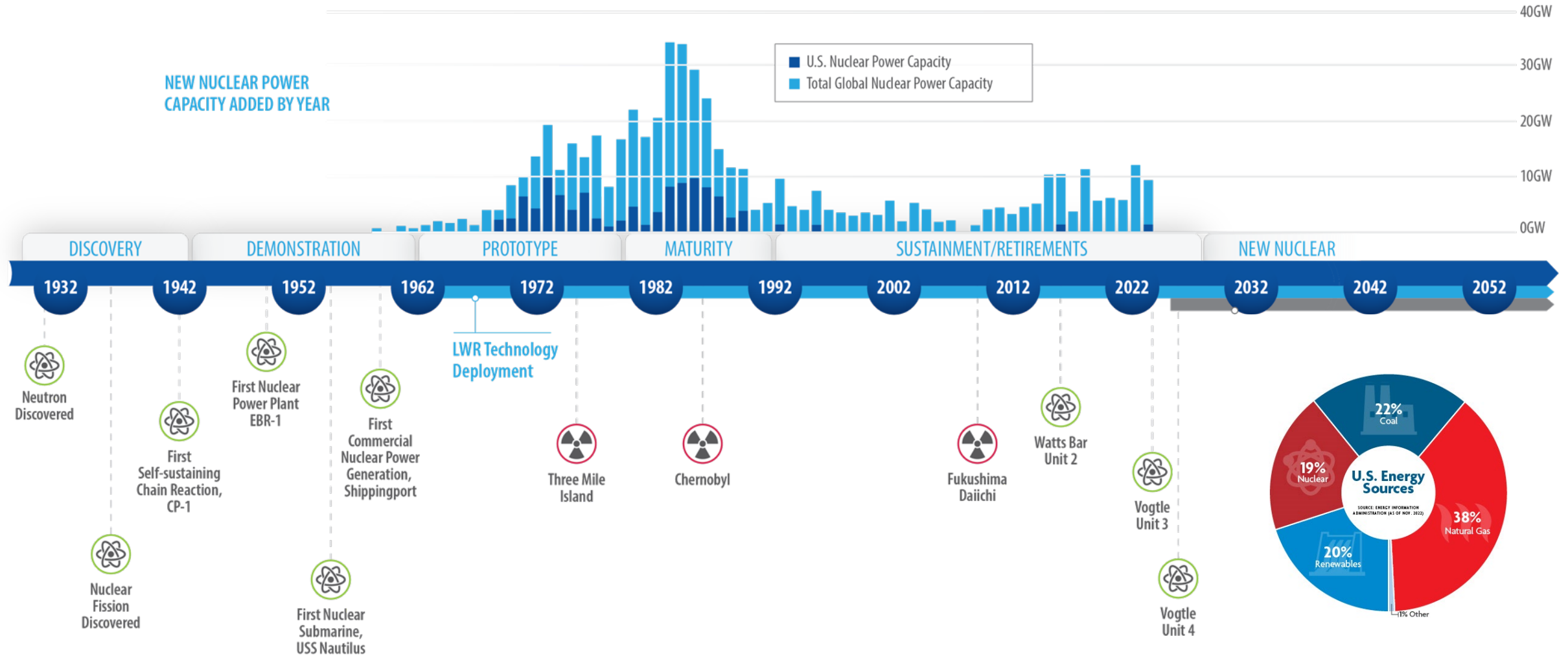
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Idaho National Laboratory

A brief history of nuclear power

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The future of nuclear power

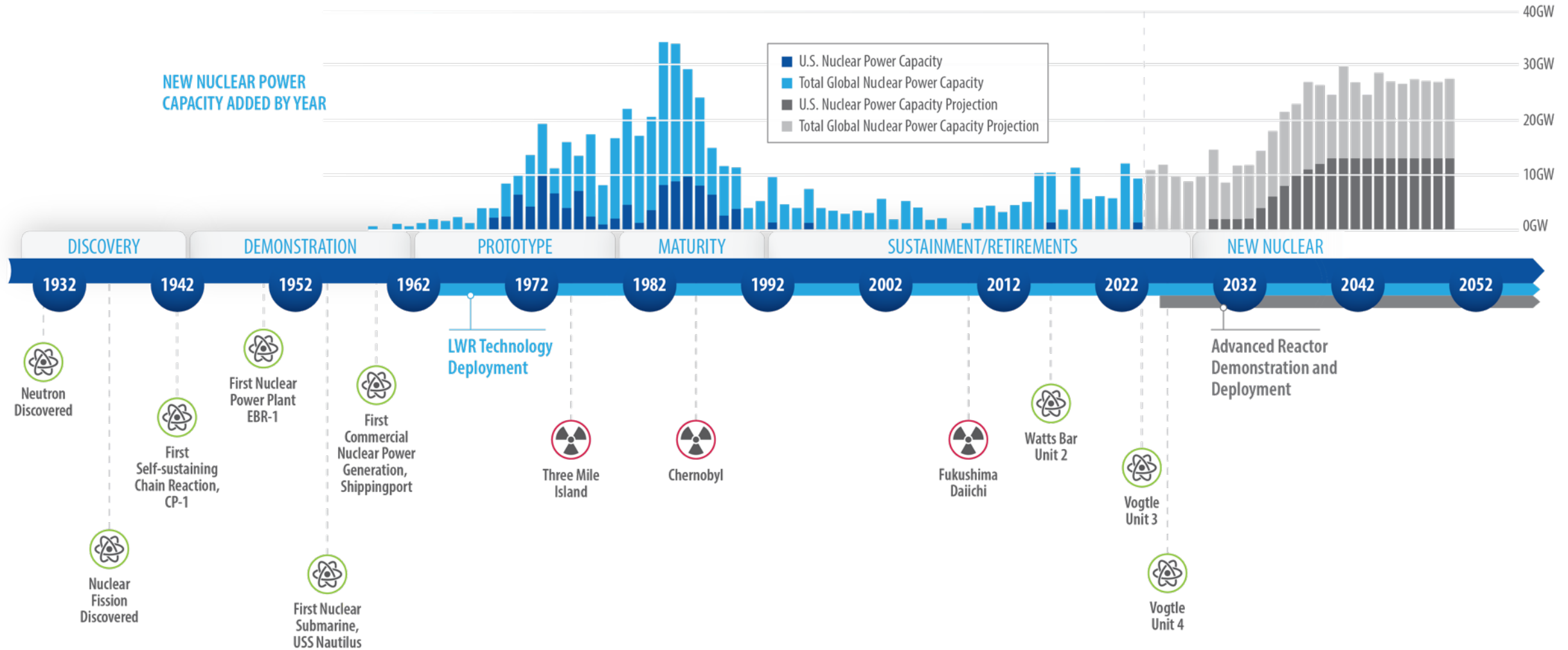
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- Countries participating in the international climate summit, the *28th Conference of the Parties (COP28)*, commit to working together to triple nuclear capacity by 2050:
 - For the United States, this would mean going from **100 GWe to 300 GWe**
 - World-wide, this would mean going from **400 GWe to 1200 GWe**
- *“Power system decarbonization modeling, regardless of level of renewables deployment, suggests that the U.S. will need ~550–770 GW of additional clean, firm capacity to reach net-zero.”*



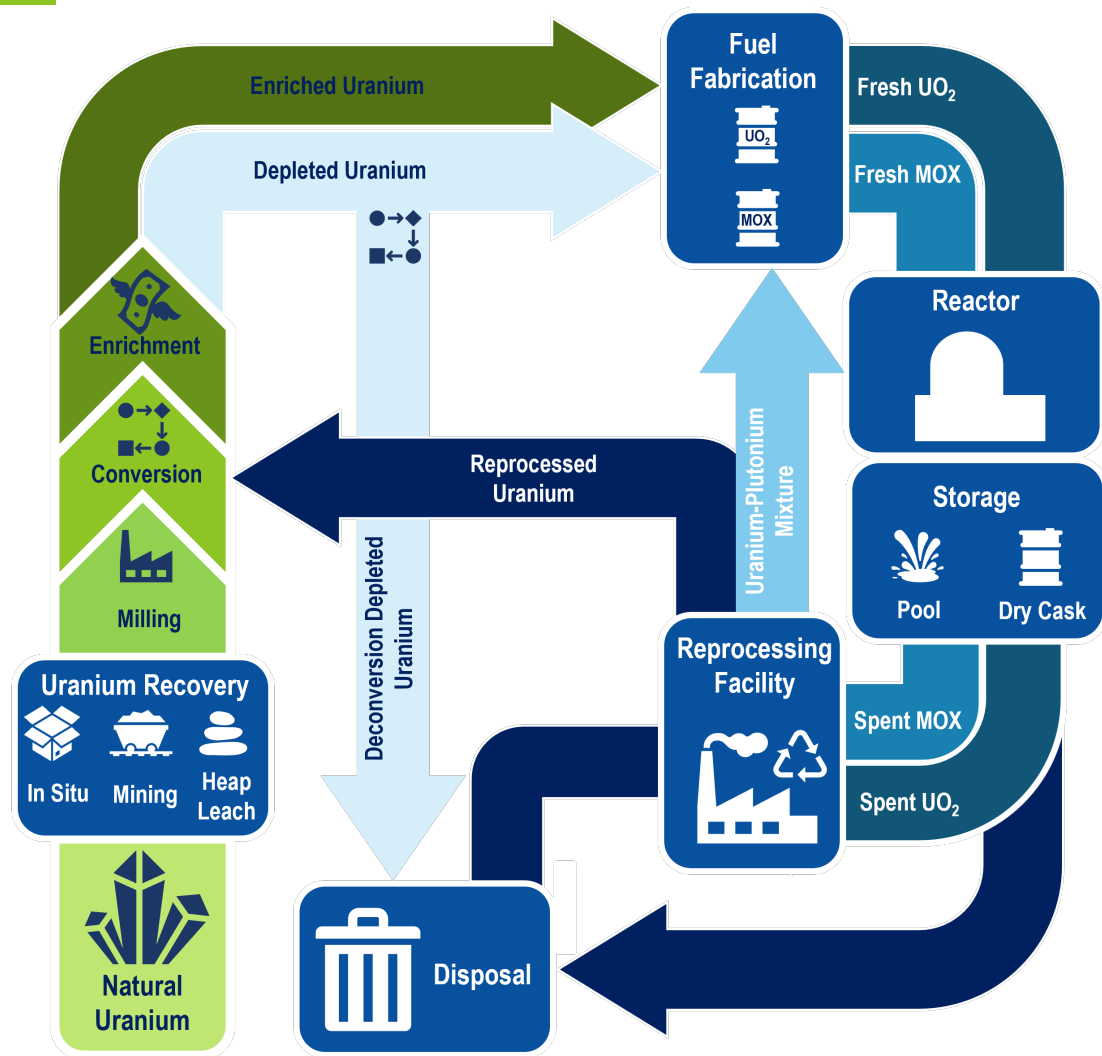
The future of nuclear power

4

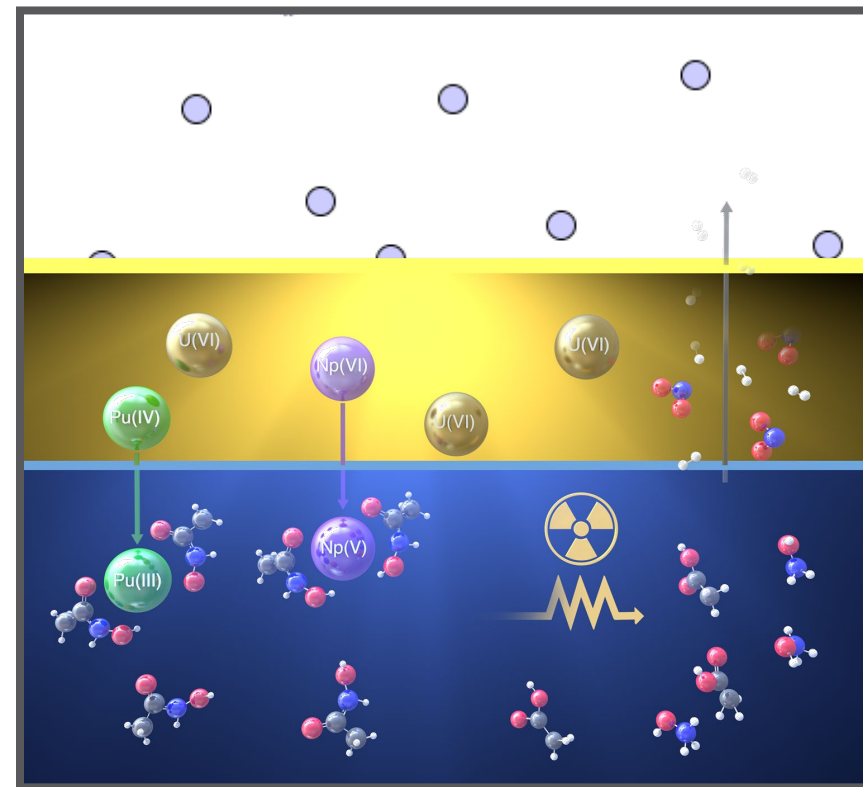


The nuclear fuel cycle + backend challenges

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A typical reprocessing solvent system is comprised of a concentrated HNO_3 aqueous phase contacted with an organic phase (organic extractants and diluent).



Reprocessing solvent system radiation chemistry

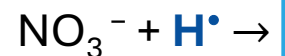
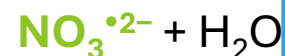
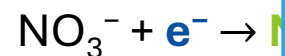
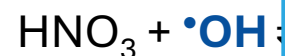
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Water Radiolysis

Direct Radiation Effects



Indirect Radiolysis



Key Radiolysis Products

e_{aq}^- , H^\bullet , OH^\bullet , and H_2O_2 from H_2O

NO_3^\bullet and HNO_2 from HNO_3

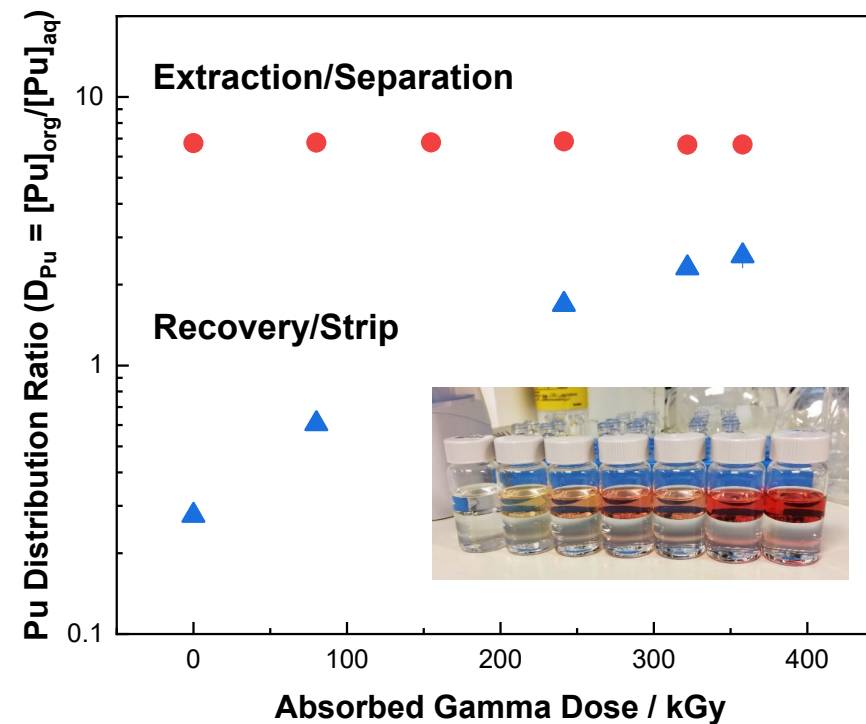
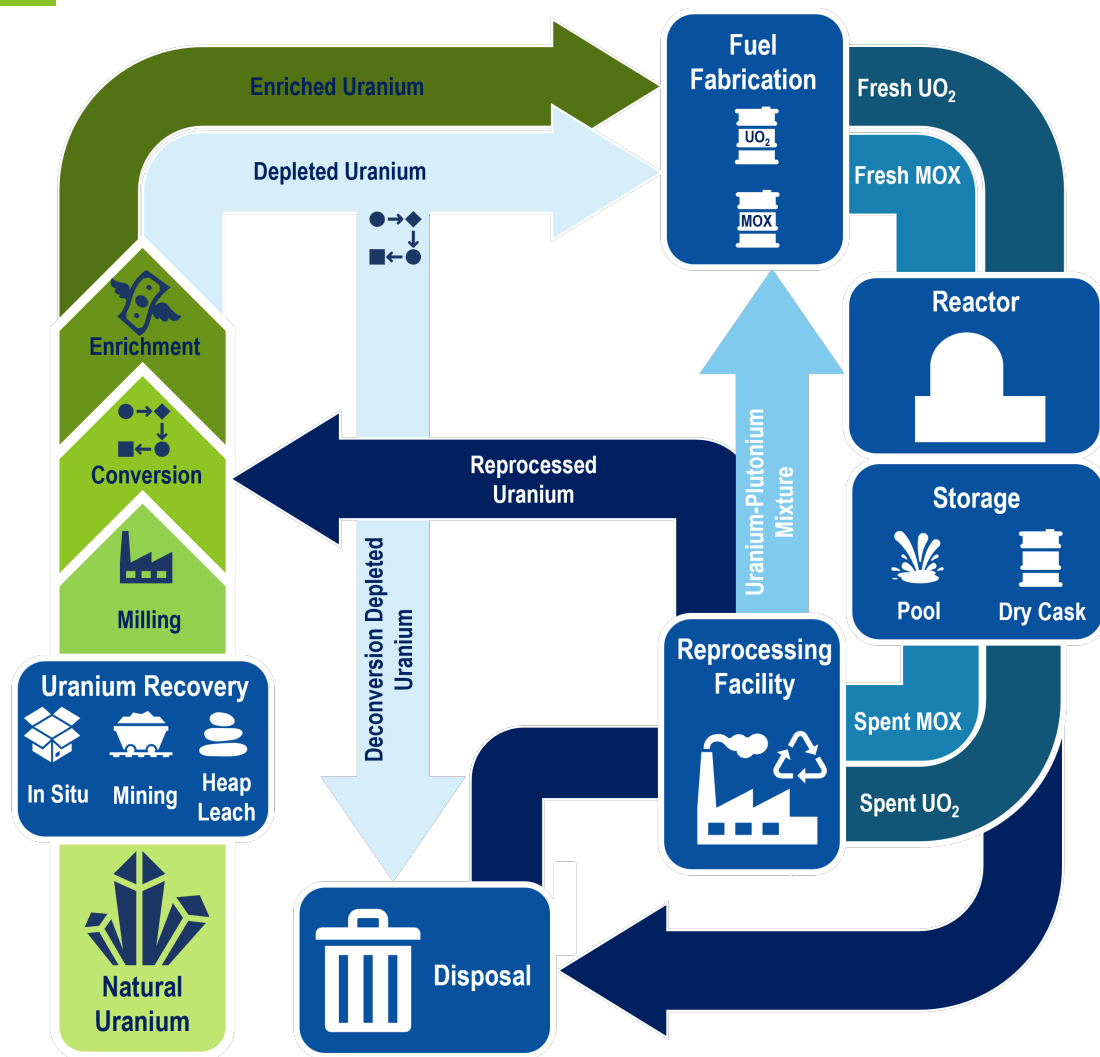
e_s^- and $\text{RH}^{\bullet+}$ from organic diluent

, H_2

- Buxton, Greenstock, Helman, and Ross, *J. Phys. Chem. Ref. Data* **1988**, 17, 513.
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Ligand radiolysis and degradation product formation

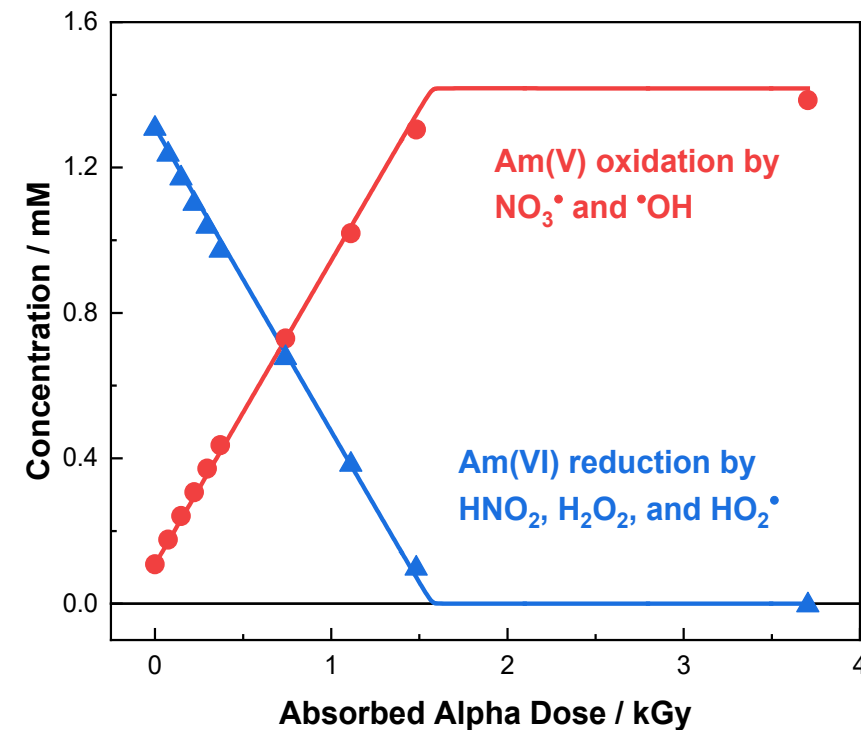
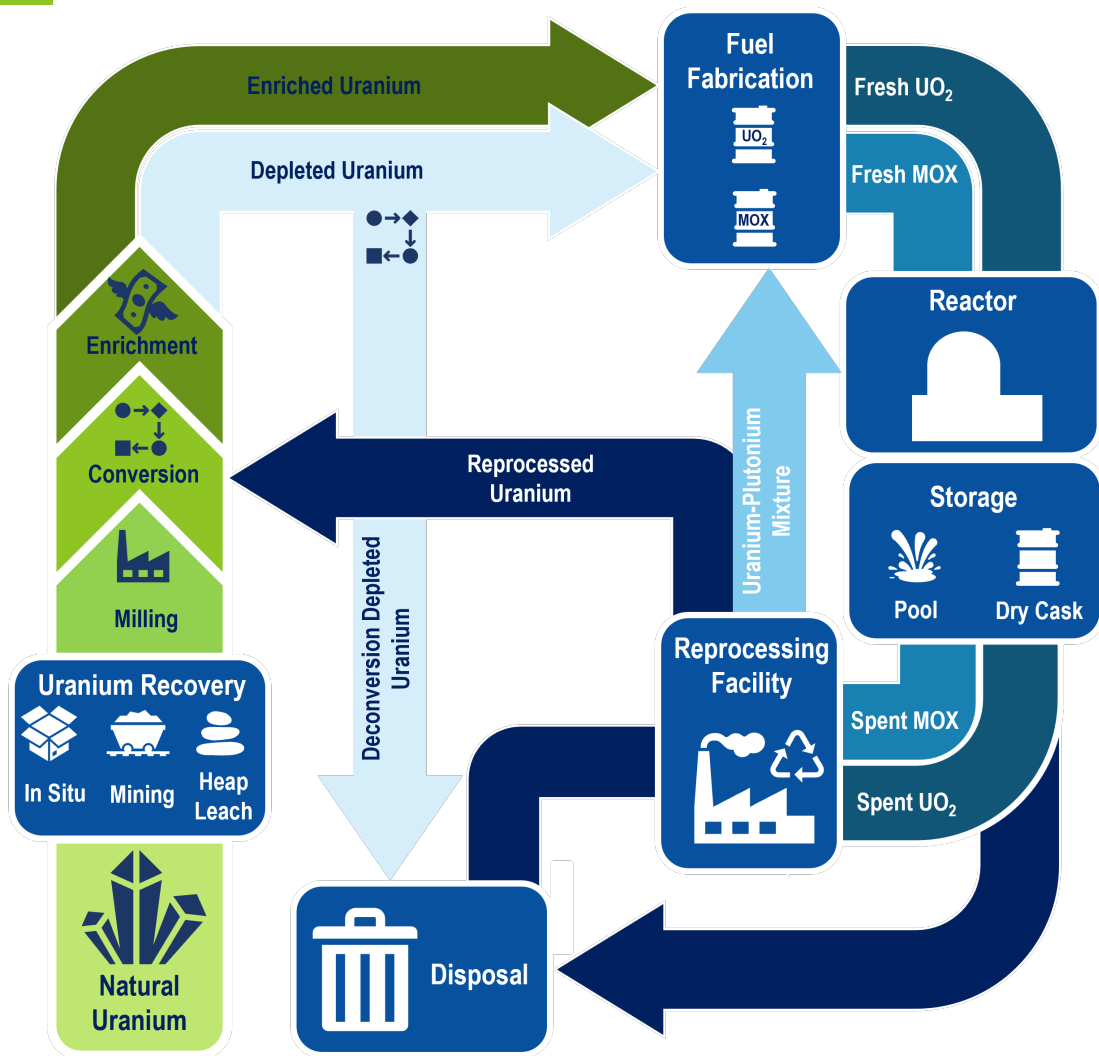
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Pu distribution ratios for extraction (▲) and strip (●) conditions as a function of absorbed gamma dose for pre-irradiated DEHBA/n-dodecane contacted with 3 M HNO₃.

Radiation-induced redox chemistry

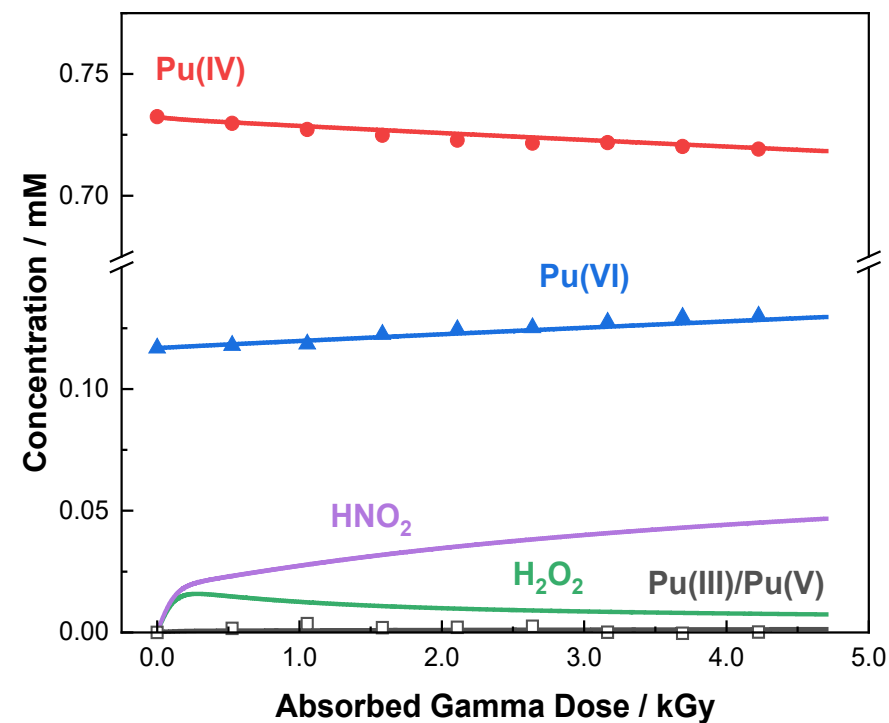
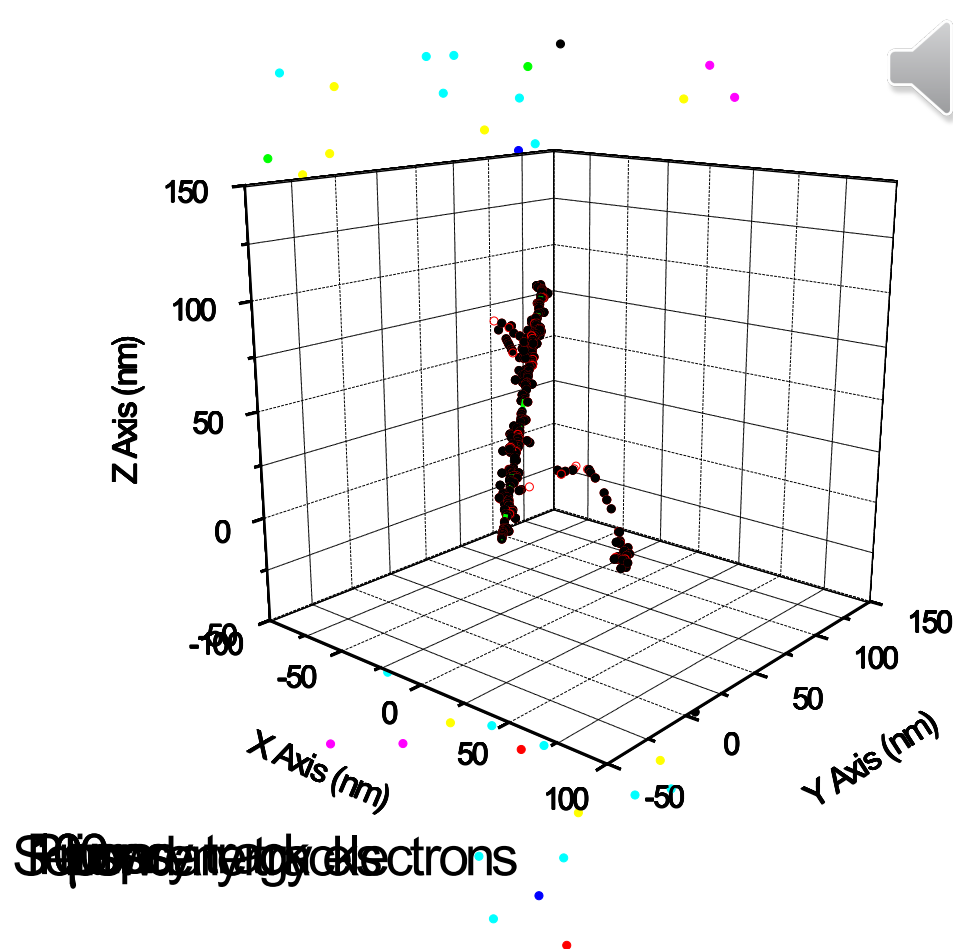
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Concentration of Am(VI) (\blacktriangle) and Am(V) (\bullet) as a function of absorbed gamma dose from the irradiation of 2 mM Am in aerated 3.0 M HNO_3 solution.

The dream: predictive models for engineer-scale processes

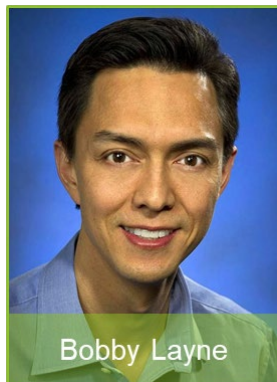
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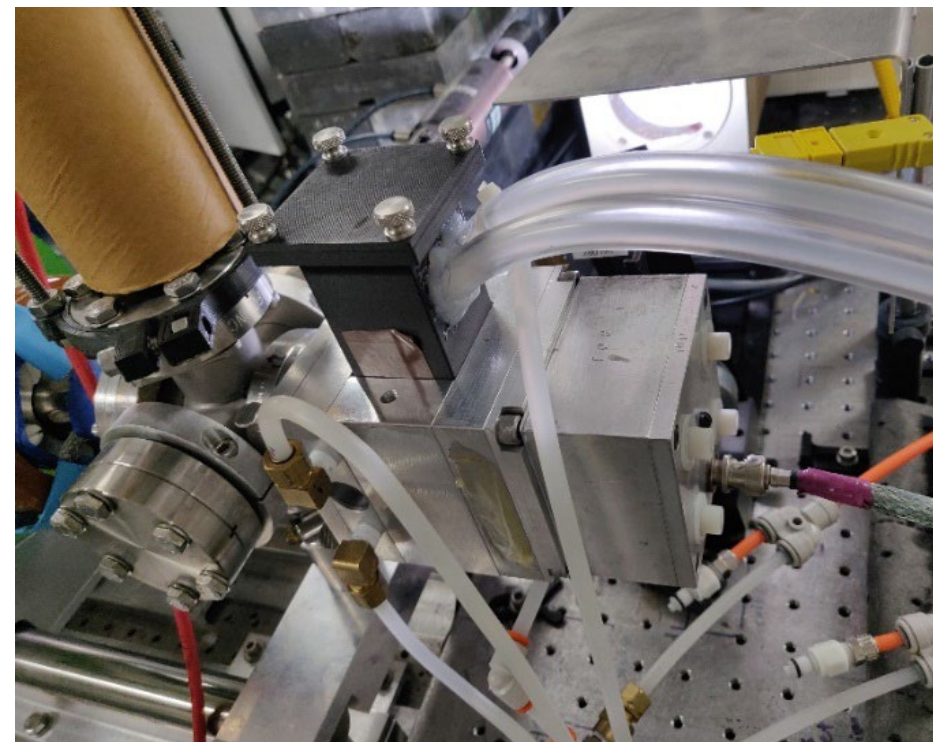
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- Pimblott, LaVerne, and Mozumder, *Journal of Physical Chemistry* **1996**, 100, 8595.
- Kynman, Grimes, Conrad, Pimblott, and [Horne](#), *Inorganic Chemistry* **2024**, 63(18), 8092.

Reaction mechanisms and activation parameters

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LEAF actinide sample holder

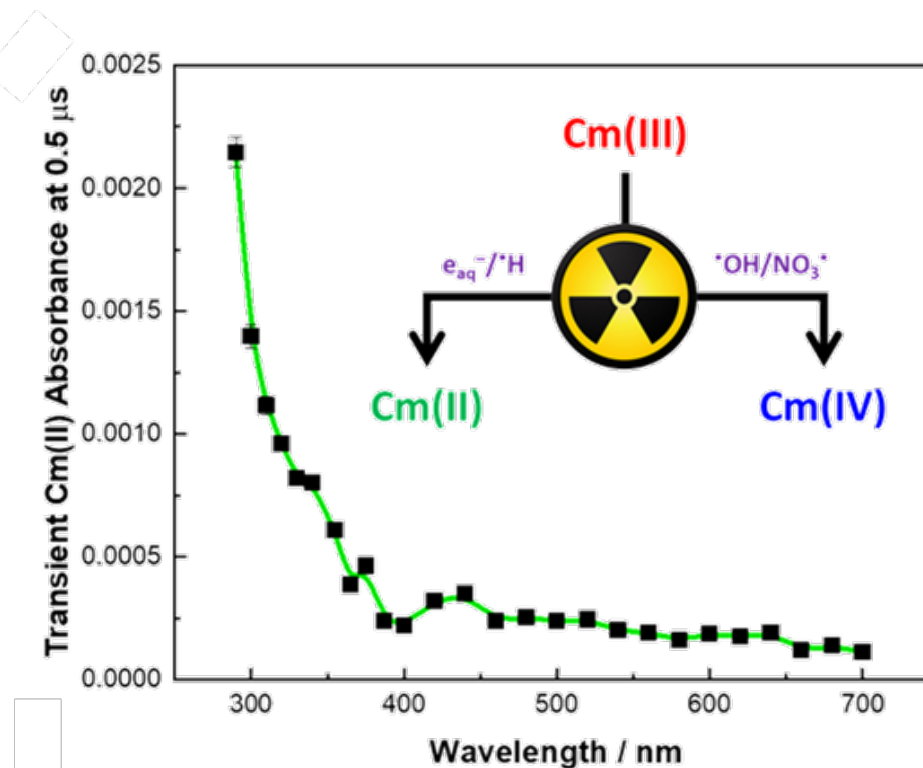
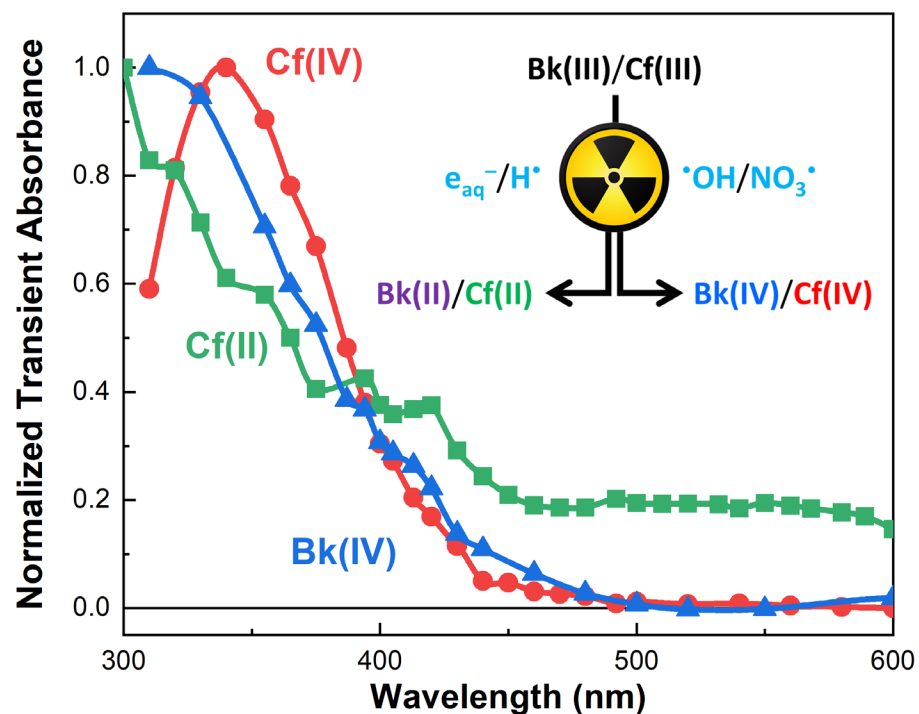


New temperature-controlled holder

- Wishart, Cook, and Miller, *Review of Scientific Instruments* **2004**, 75(11), 4359.
- **Horne**, Grimes, Zalupski, Meeker, Albrecht-Schönzart, Cook, and Mezyk, *Dalton Transactions* **2021**, 50, 10853.
- **Horne**, Rotermund, Grimes, Sperling, Meeker, Zalupski, Beck, Gomez Martinez et al., *Inorganic Chemistry* **2022**, 61(28), 10822.
- Rotermund, Mezyk, Sperling, Beck, Wineinger, Cook, Albrecht-Schönzart, and **Horne**, *Journal Physical Chemistry A* **2024**, 128(3), 590.
- Kynman, Grimes, Mezyk, Layne, Cook, Rotermund, and **Horne**, *Dalton Transactions* **2024**, 53, 9262.

Radiation-induced actinide redox chemistry

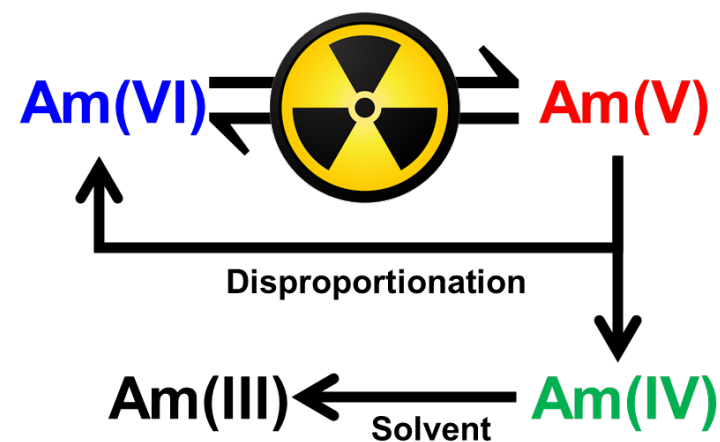
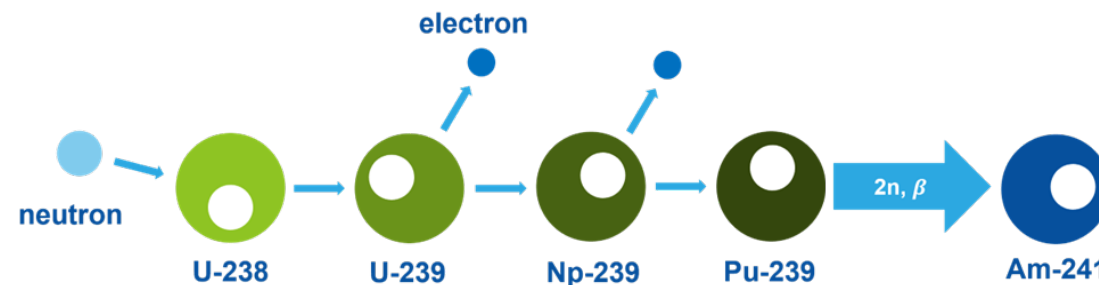
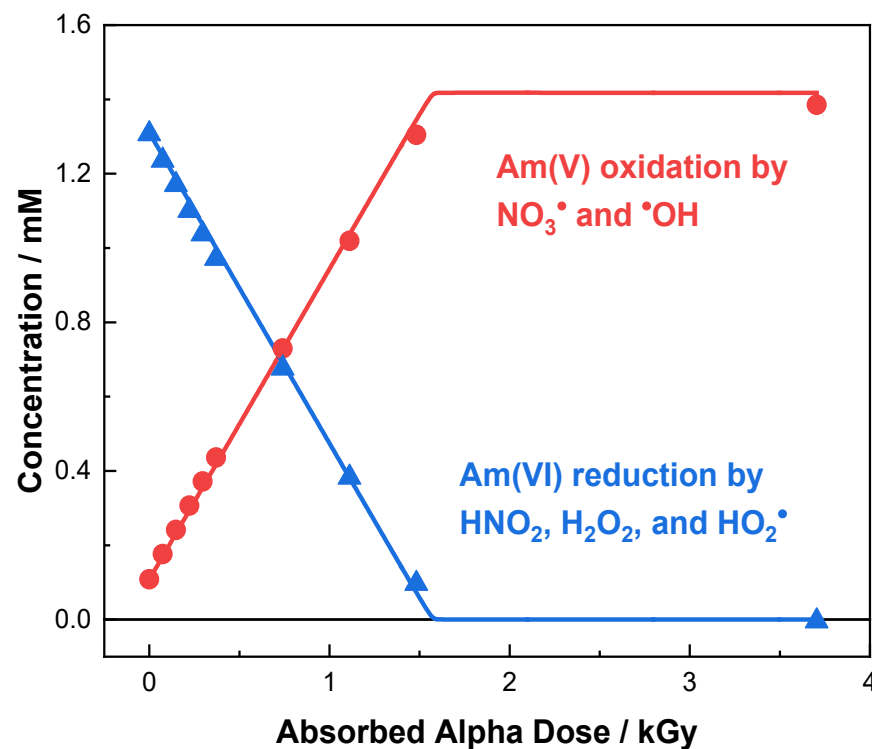
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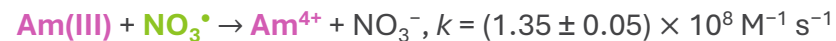
Radiation-induced americium redox chemistry

12

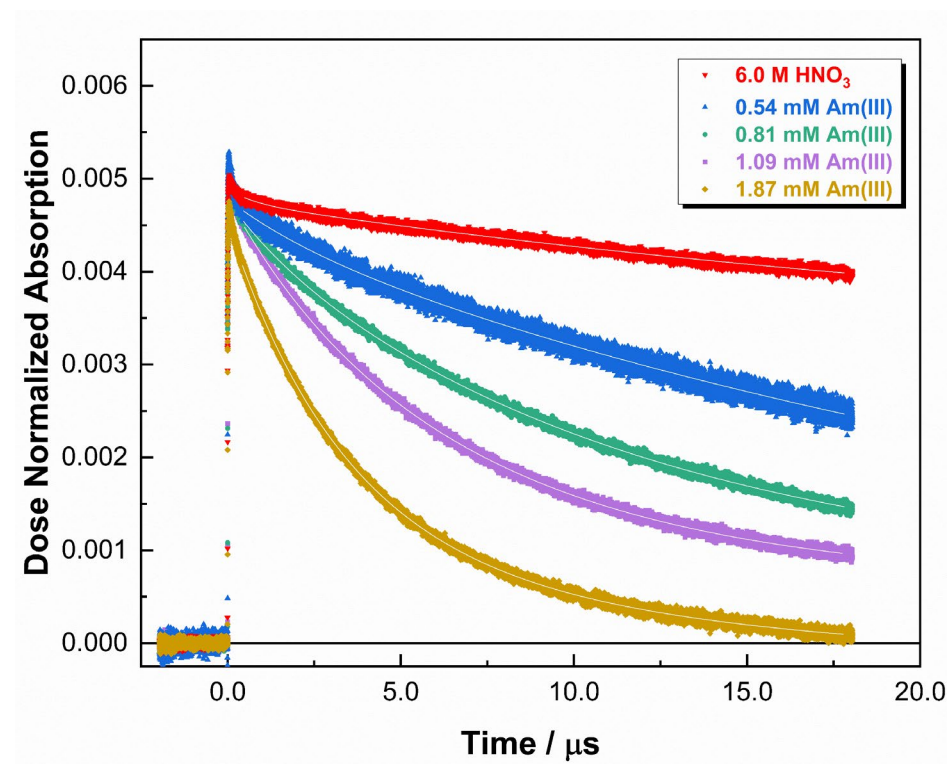
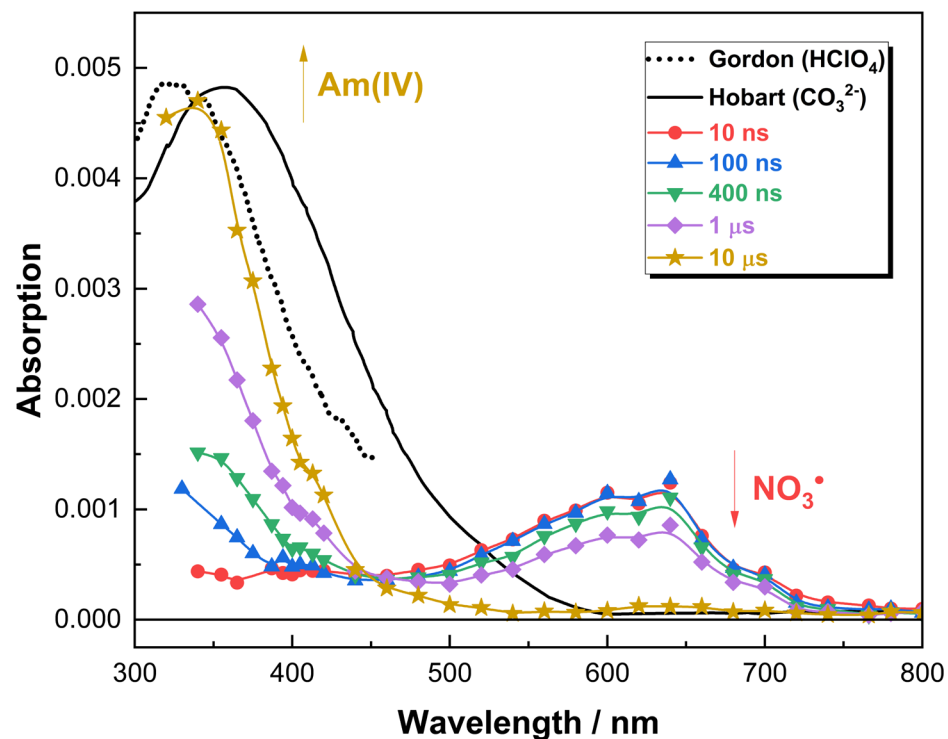


Radiation-induced americium(IV) redox chemistry

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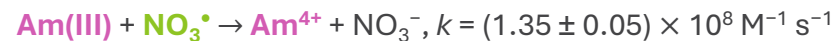


Am(IV) lifetime in 6 M HNO₃ is approximately 16 μs

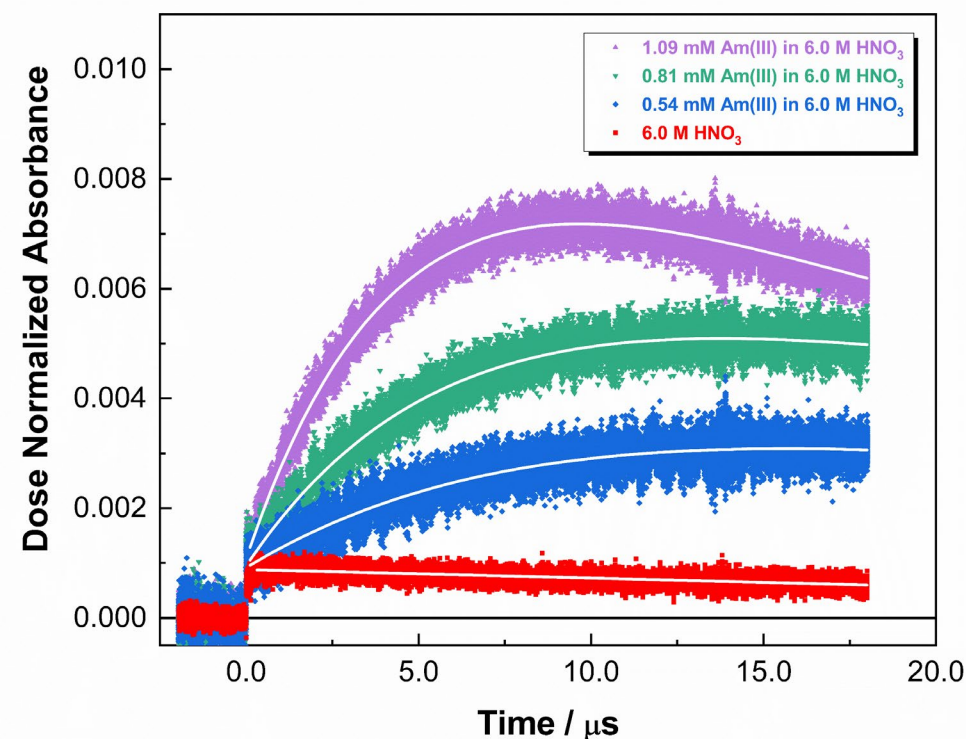
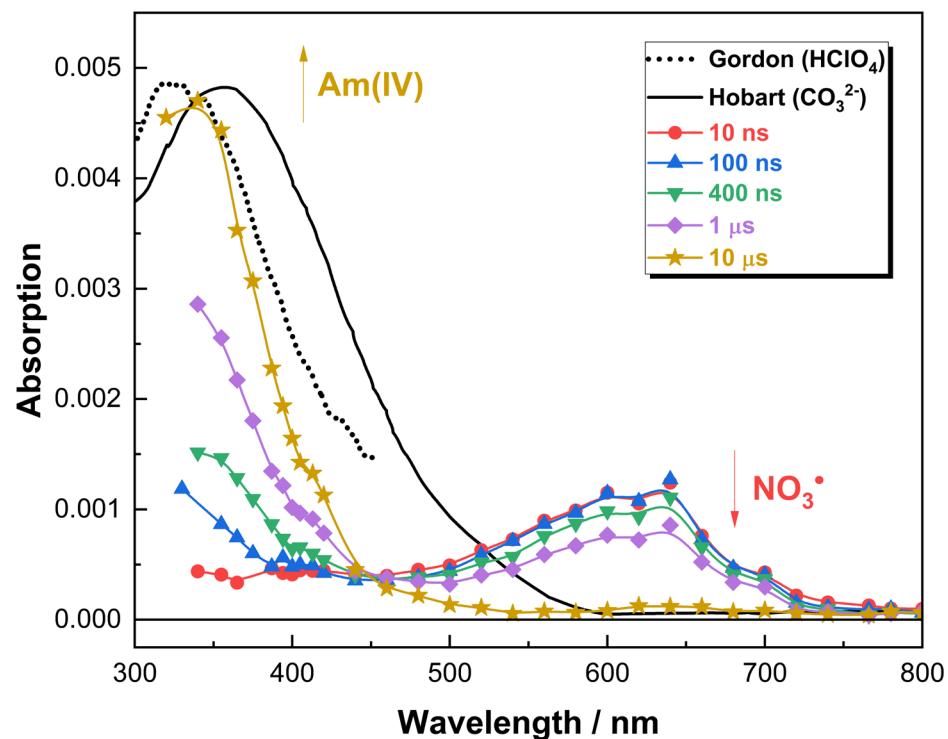


Radiation-induced americium(IV) redox chemistry

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Am(IV) lifetime in 6 M HNO_3 is approximately 16 μs

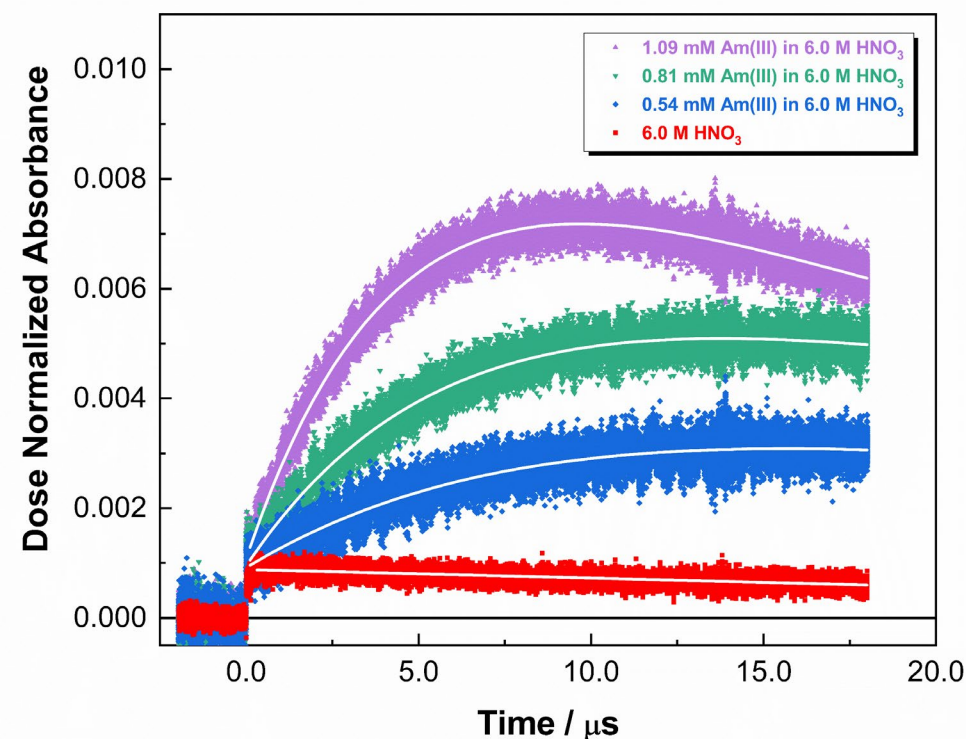
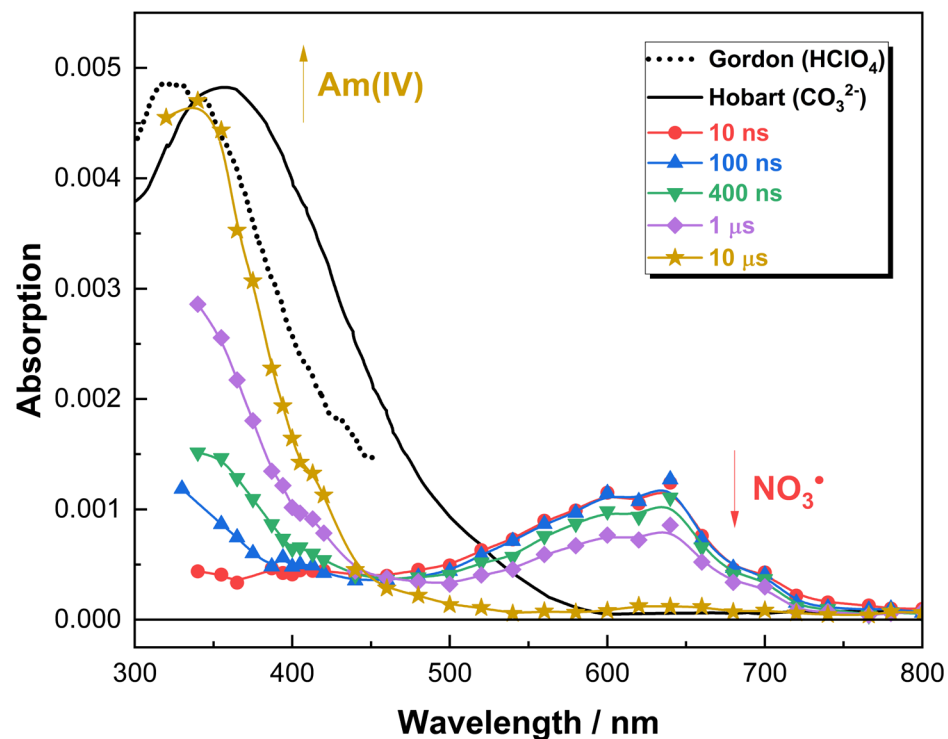


Radiation-induced americium(IV) redox chemistry

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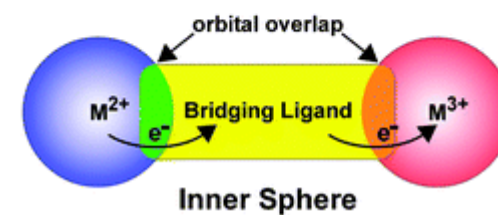
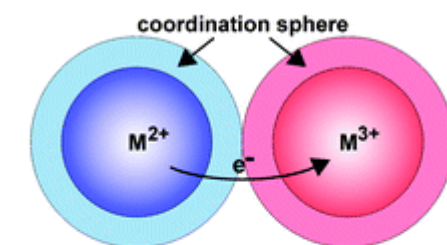
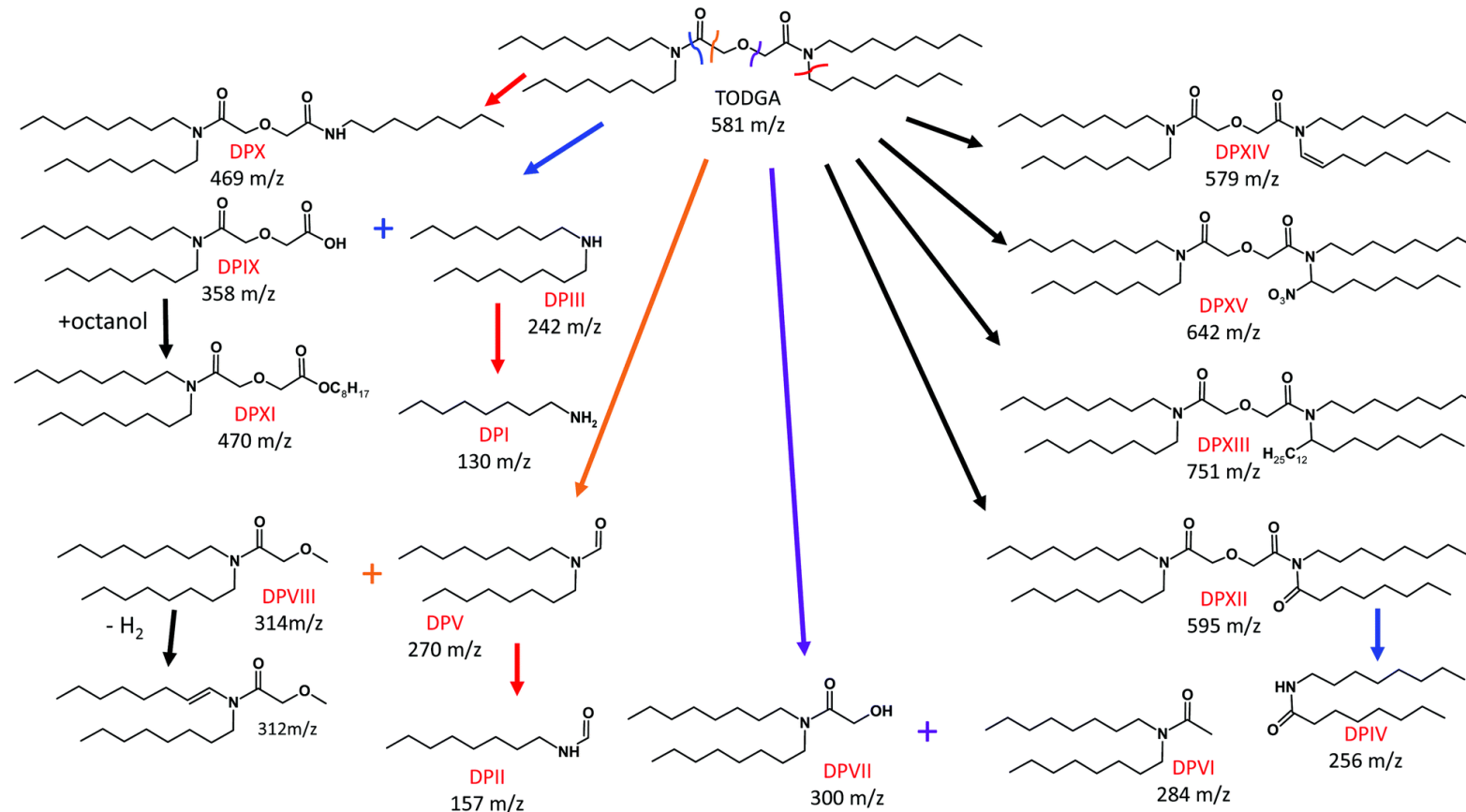
Arrhenius: $E_a = 19.5 \text{ kJ mol}^{-1}$ and $A = 4.2 \times 10^{11} \text{ s}^{-1}$

Eyring: $\Delta H^\ddagger = 17 \text{ kJ mol}^{-1}$ and $\Delta S^\ddagger = -30.7 \text{ J mol}^{-1} \text{ K}^{-1}$

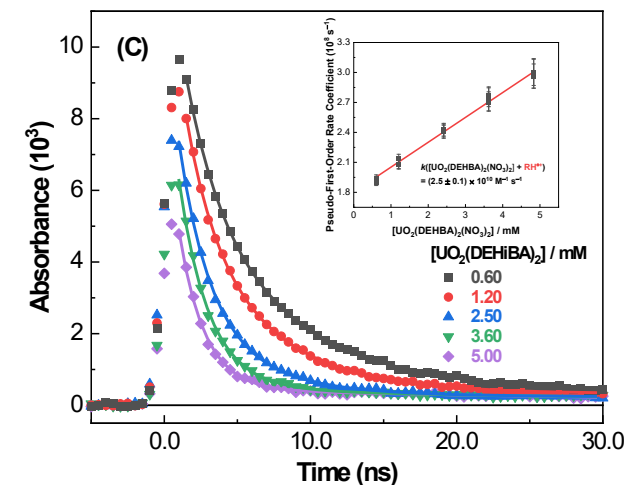
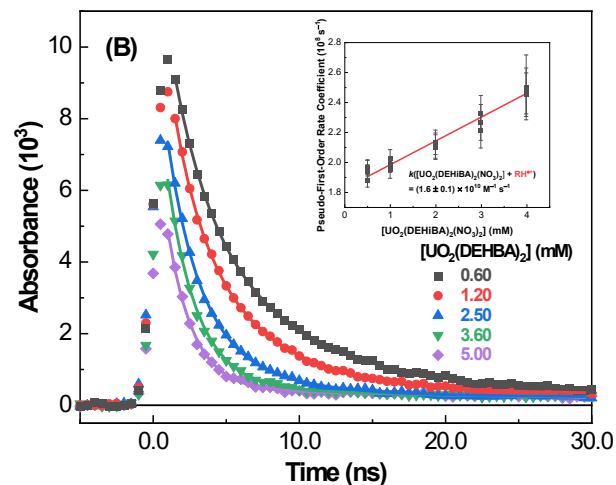
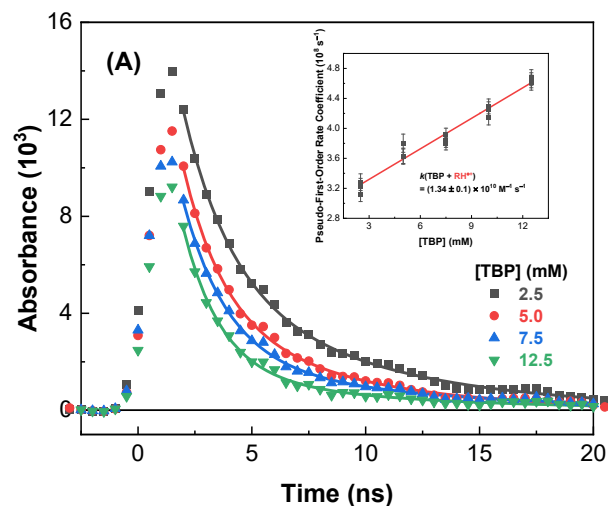
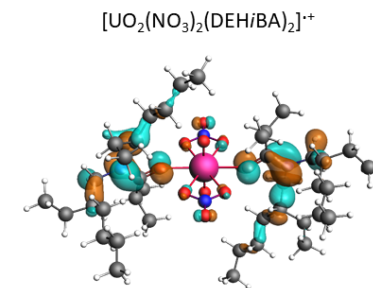
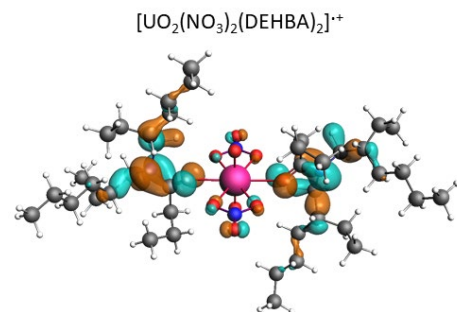
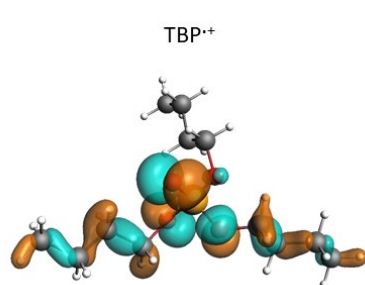


What about metal ion complexation effects?

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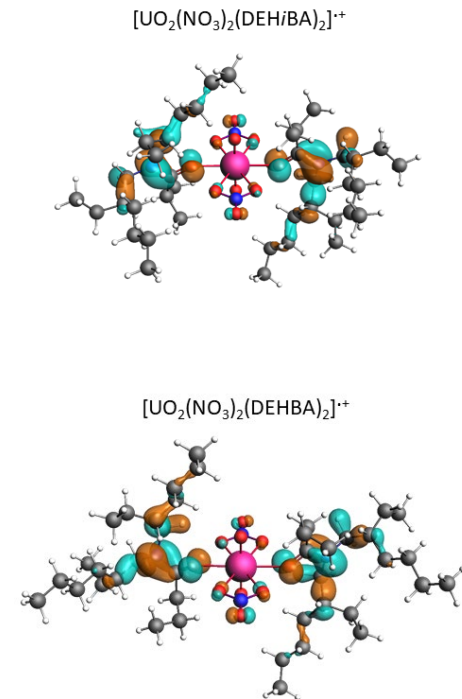
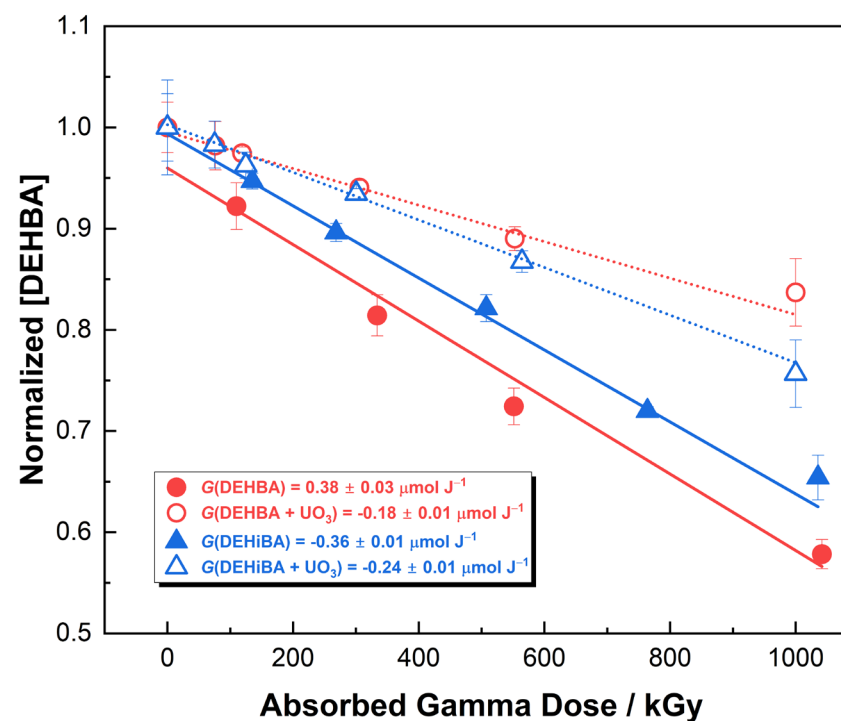
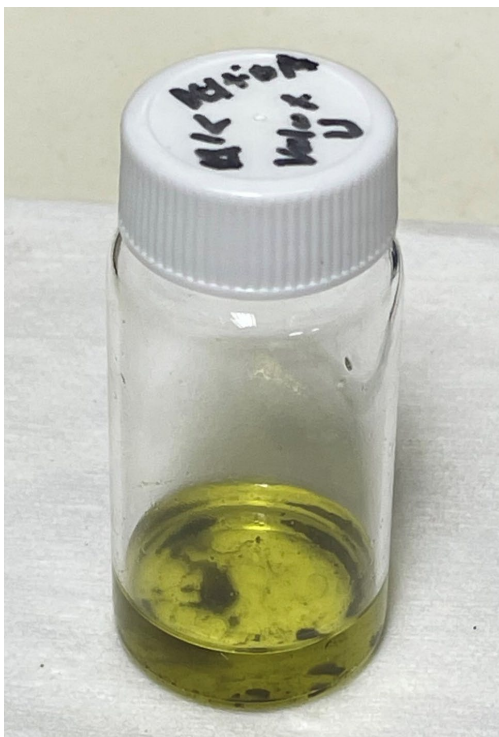
Uranium complexation effects



- **U(VI)** complexation had negligible effect on the reaction of **TBP** with **RH⁺⁺**, $k = (1.3 \pm 0.1) \times 10^{10} \text{ M}^{-1} \text{ s}^{-1}$.
- For **DEHBA** and **DEH/BA**, **U(VI)** complexation afforded a 2.6× and 1.4× increase in their respective rate coefficients.

Gamma irradiation of uranium loaded solvent systems

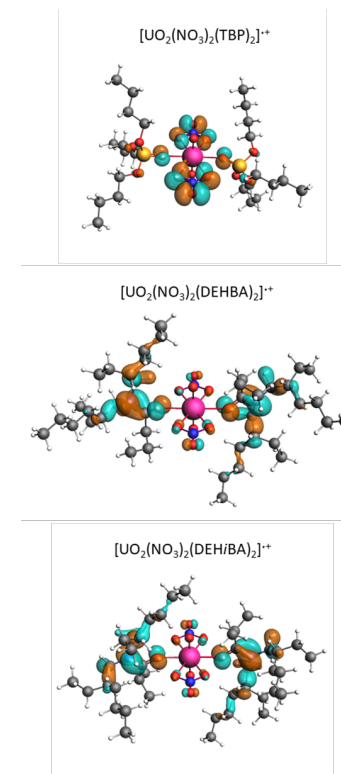
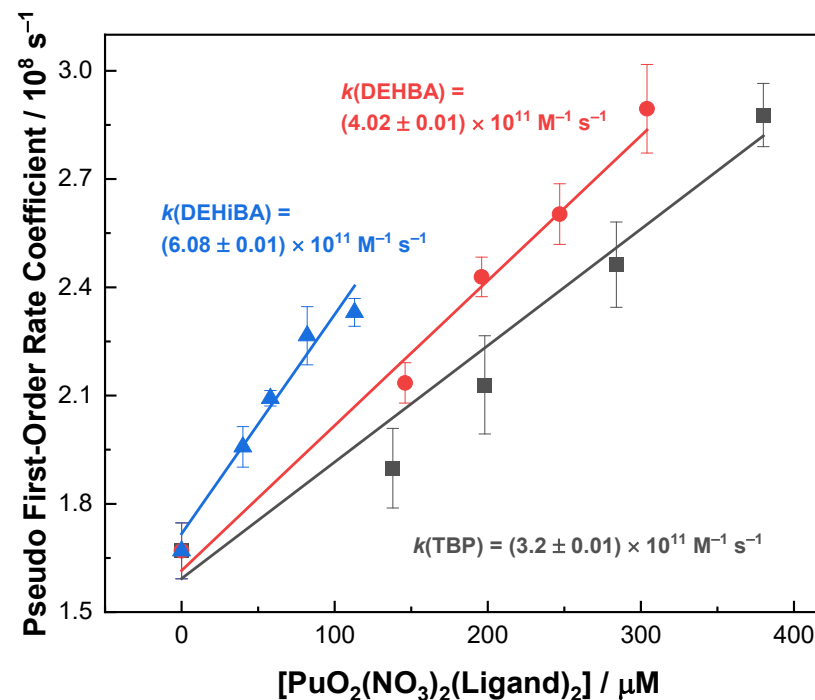
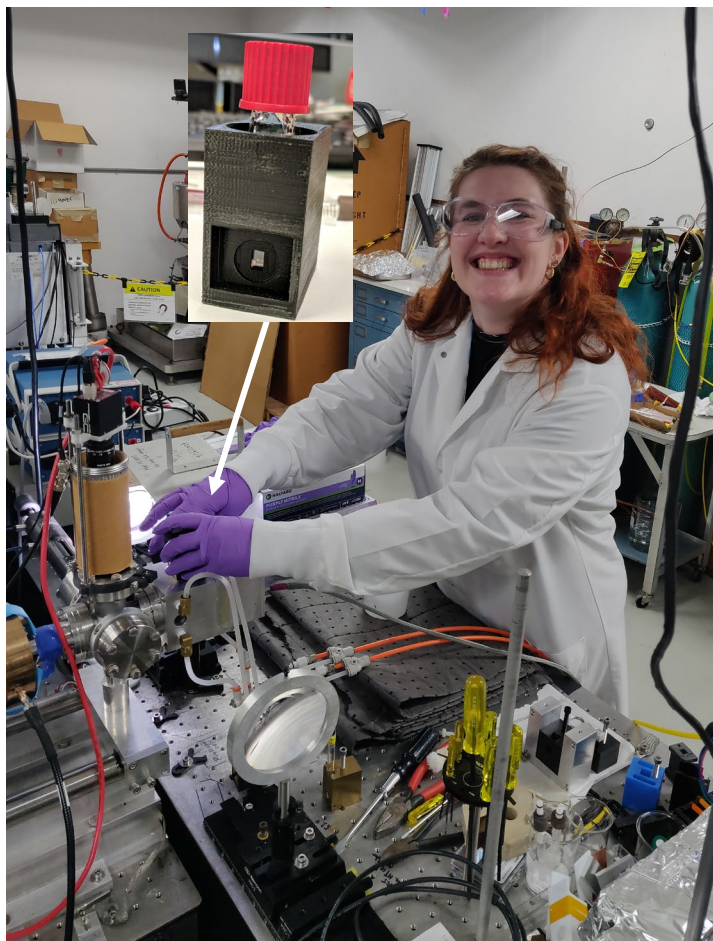
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- Normalized loss of **DEHBA** (●/○) and **DEHiBA** (▲/△) in deaerated, 6.0 M HNO₃ pre-equilibrated *n*-dodecane solution in the presence and absence of ϵ -UO₃ as a function of absorbed cobalt-60 gamma dose.

Electron pulse irradiation of plutonium loaded solvent systems

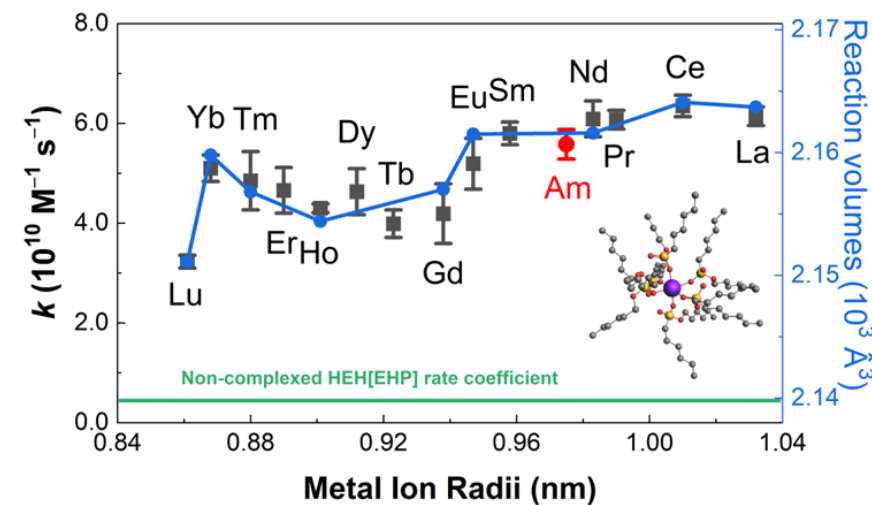
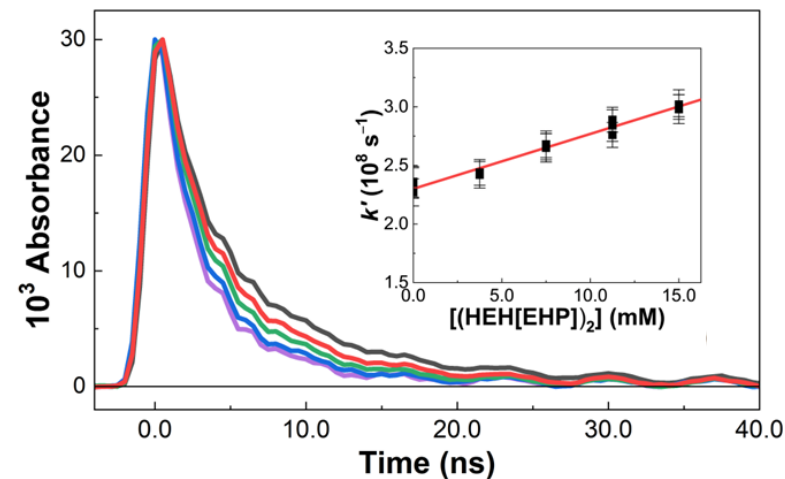
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- **Pu(VI)** complexation afforded significantly faster rates for the reaction of RH^{+} with $[\text{PuO}_2(\text{NO}_3)_2(\text{L})_2]$ as compared to the non-complexed ligands ($\sim 10^{10} \text{ M}^{-1} \text{ s}^{-1}$).

Overview

- Radiation-induced processes have the capacity to drive backend nuclear fuel cycle systems far from equilibrium, impacting their performance and longevity.
- Electron pulse radiolysis is an essential technique for elucidating complex reaction mechanisms, especially those involving short-lived transients.
- Metal ion complexation has profound effect on the chemical kinetics and radiation robustness of organic ligands.



Acknowledgements

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Amy Kynman



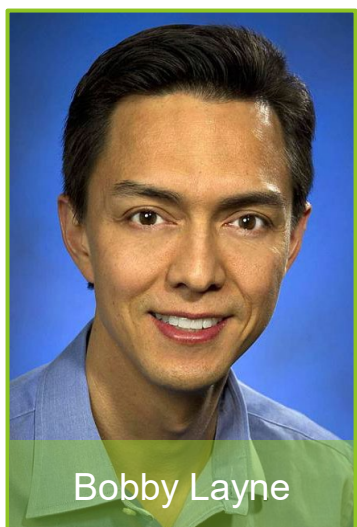
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